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Solubility

*UNIT 3 PPA 1*

**INTRODUCTION**

Compounds called fertilisers are added to the soil to help plants to grow well. A compound can be used as a fertiliser if it contains one of the essential elements, nitrogen (N), phosphorus (P) or potassium (K) and if it dissolves in water.

Ammonium compounds (N), potassium compounds (K), nitrate compounds (N) and phosphate compounds (P) are all possible fertilisers because each one contains an essential element.

However, only the ones which are soluble in water will be able to fertilise plants.

The aim of this experiment is to test the solubility in water of some ammonium, potassium, nitrate and phosphate compounds in order to decide if they could be used as fertilisers.

**You will need**

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| test tubes & rack | beaker |
| Samples of solid: | ammonium sulphate, potassium nitrate, sodium nitrate, calcium phosphate, ammonium phosphate |

**Safety**

The dilute solutions are of no significant hazard. Some of the solids are irrritants. Eye protection should be worn.

Diagram

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1. Add dilute hydrochloric acid to the beaker until it is half full.

Shape

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1. Pour water into a test tube to a depth of 3 - 4 cm.
2. Shape

   Description automatically generated with low confidenceUsing a spatula add a tiny amount (about the size of half a pea) of ammonium sulphate to the water.

Shape, arrow

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1. Hold the test tube at the mouth and 'flick' it back and forth for several minutes.
2. Look at the mixture to see if the solid has dissolved.
3. Record the result in the table by writing down the name of the compound and whether it was soluble or insoluble.
4. Repeat the experiment with each of the remaining compounds recording the result each time.

**Results sheet**

*What was the aim of the experiment? (Remember to say which factor you were investigating)*

**Procedure**

*Describe briefly how you tested the solubility of the compounds.*

*Complete the following table:*

*Table

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**Conclusion**

*Which compound could not be used as a fertiliser?*